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1 a) The relative atomic mass of an element is the weighted average mass of the isotopes of the element. (1) It is measured on a scale on which a carbon-12 atom has a mass of exactly 12 units. (1).  
b) The total mass of 100 atoms =  $(75 \times 35) + (25 \times 37) = 3550$  (1) The average mass of 1 atom =  $3550/100 = 35.5$  (1)

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