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Assuming the pressure stays constant, at what temperature will the volume of the gas be 604 cm³? Given: V₁ = 506 cm³. T₁ = 147 °C. V₂ = 604 cm³. T₂ = ? V₁/T₁ = V₂/T₂. 506 cm³/420 K = 604 cm³/T₂. T₂ = 501 K . or. 501 - 273 = 228 °C. Remember to convert temperature to Kelvin (420 K) Episode 902

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