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Stoichiometry Study

Chem chapter 11 vocab

Stoichiometry: The study
of quantitative

relationships between the

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amounts of reactants used and products formed by a chemical reaction; is based on the law of conservation of mass. mole ratio: In a balanced equation, the ratio between the numbers of moles of any two substances. limiting reactant: A reactant that is totally consumed during a chemical reaction, limits the ...

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Stoichiometry. the study
of the Quantitative, or
measurable relationships

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that exist in chemical formulas and chemical reactions. the Law of Conservation of mass. Stoichiometry is based on _____

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stoichiometry. mole
ratio. limiting reactant.
excess reactant. the study
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relationships between the amounts of.... in a balanced equation, the ratio between the number of moles.... a reactant that is totally consumed during a chemical reaction....

Stoichiometry Chapter
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Key

368 Chapter 11 •

Stoichiometry Section

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11.11.1 Objectives

Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation.

Review Vocabulary

reactant: the starting substance in a chemical reaction

New Vocabulary

stoichiometry mole ratio

Defining Stoichiometry

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Chapter 11:
Stoichiometry

11.1 Defining

Stoichiometry 11.2

Stoichiometric

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Defining Stoichiometry

11.2 ...

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Stoichiometry ...

In Section 11.3, for

example, you learned

how to express the

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stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water).

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209 Stoichiometry

Stoichiometry CHAPTER 11

SOLUTIONS

MANUAL Section 11.1

Defining Stoichiometry

pages 368 – 372 Practice

Problems pages

371 – 372 1. Interpret the

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following balanced
chemical equations in
terms of particles, moles,
and mass. Show that the
law of conservation of
mass is

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Key Stoichiometry is the
tool for answering these
questions. Stoichiometry
The study of quantitative

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relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

Chapter 11

Stoichiometry Study

Answer Key

Chapter 11 187 Exercise

11.3 – Equation

Stoichiometry: Iron is combined with carbon in

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a series of reactions to form pig iron, which is about 4.3% carbon.

$$2\text{C} + \text{O}_2 \rightarrow 2\text{CO}$$
$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$

2Fe 3CO₂ 2CO C (in iron) CO₂ Pig iron is easier to shape than pure iron, and the presence of carbon lowers its melting point

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Number of Atoms •

Number of Molecules •

Number of Moles •

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Chapter 11 Stoichiometry

The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry.

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Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles

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