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1) calculate moles of each element (if given mass) 2) divide each mole quantity by the smallest mole value from step 1. 3) use whole number values from step 2 as the subscripts to the original equation to write the

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~~empirical formula. Percent Composition. -the relative amounts of each element in a compound.~~

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Multiple Choice 24) What is the amount of heat required to raise the temperature of 200.0 g of aluminum by 10°C? cal (specific heat of aluminum = 0.21 J/g x °C 24)  
A) 4200 cal B) 420 cal C) 42 000 cal D) 420 000 cal E) none of the above

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statement or answers the  
question. \_\_\_\_ 1. How many  
moles of tungsten atoms are  
in  $4.8 \times 10^{25}$  atoms of

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c.  $1.3 \times 10^{?1}$  moles b.  $8.0 \times$   
 $10^1$  moles d.  $1.3 \times 10^{?2}$   
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1600s - Robert Boyle studied the behavior of gases and is responsible for the theory of elements ...

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