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Worksheet 1 Balancing

Balancing Chemical

Equations Answer Key

Balance the equations

below: 1) $1 \text{ N}_2 + 3 \text{ H}_2$

$\rightarrow 2 \text{ NH}_3$ 2) 2 KClO_3

$\rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) 2

$\text{NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} +$

1 Cl_2 4) $2 \text{ H}_2 + 1 \text{ O}_2$

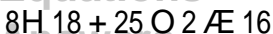
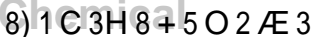
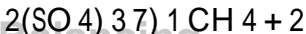
$\rightarrow 2 \text{ H}_2\text{O}$ 5) 1 Pb(OH)

$2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1$

PbCl_2 6) $2 \text{ AlBr}_3 + 3 \text{ K}$

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Equations

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Equations Balance the

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equations below: 1) _____



NH₃ ... 7) Copper and sulfuric acid react to form copper (II) sulfate and water and sulfur dioxide.

8) Hydrogen gas and nitrogen monoxide react to form water and nitrogen gas.

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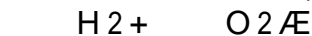
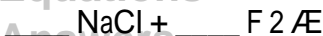
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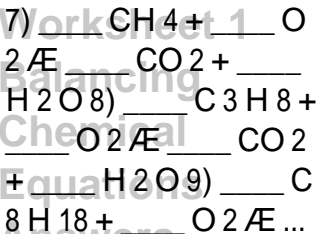
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+ Equations

Answers

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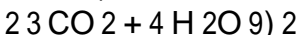
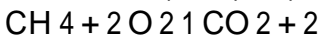
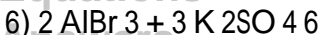
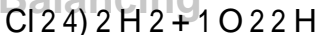
Key Balance the

equations below: 1) 1 N_2

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H₂SO₄ NaNO₃ HNO₃ Na₂SO₄

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in balancing chemical
equations, follow the tips
for correct balancing
chemical equations
worksheet answers. Tip #
1: When you are trying to
balance the chemical
equations, you should

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remember that you can only change the value of coefficient in front of the element or compound, and not the subscript.

Answers

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described in section 4.1 is balanced, meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides. This is a requirement the equation must satisfy to be consistent with the law of

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Equations chemical
reaction that absorbs
energy from its
surroundings. Law of
Conservation of Energy
energy cannot be created
or destroyed, but it can
be transferred or
transformed from one
form to another. Chapter

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Worksheet 1 Answers

As illustrated in Figure

7.3.2, the coefficients

allow us to interpret

Equation 7.3.1 in any of
the following ways: Two

NH_4^+ ions and one Cr

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2 O₇ 2 – ion yield 1
formula unit of Cr₂ O₃,
1 N₂ molecule, and 4 H
2 O molecules. One mole
of (NH₄)₂ Cr₂ O₇
yields 1 mol of Cr₂ O₃,
1 mol of N₂, and 4 mol
of H₂ O.

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